

exploring science fizzy metals 2 answers Handbook

Exploring Science Fizzy Metals 2 Answers: An In-Depth Guide

Exploring science fizzy metals 2 answers is an engaging topic often encountered in educational settings, especially in science classes aimed at young learners. This concept involves understanding the chemical reactions that produce fizzy or bubbling effects when certain metals interact with acids or other substances. The activity not only sparks curiosity but also provides a hands-on way to explore fundamental chemical principles such as reactivity, acids and bases, and gas production. In this article, we will delve into the details of fizzy metals, explore common questions surrounding them, and provide comprehensive answers to enhance your understanding of this fascinating subject.

Understanding the Concept of Fizzy Metals

What Are Fizzy Metals?

Fizzy metals refer to metals that produce bubbles or fizzing reactions when they come into contact with acids or other reactive substances. This fizzing is typically caused by the release of gas, most commonly hydrogen gas, during a chemical reaction. Common examples of fizzy metals include zinc, magnesium, and aluminum, which are often used in educational experiments to demonstrate chemical reactivity.

The Science Behind Fizzy Reactions

The fizzing occurs due to a chemical process called *metal-acid reaction*. When a metal reacts with an acid, it displaces hydrogen atoms from the acid molecules, resulting in the formation of a metal salt and hydrogen gas. The hydrogen gas then escapes as bubbles, creating a fizzy or bubbling appearance. The general reaction can be summarized as:



For example, when zinc reacts with hydrochloric acid:



Common Questions and Answers About Fizzy Metals

1. Why do some metals produce more fizz than others?

The amount of fizz produced depends on the metal's reactivity with acids. Metals higher in the reactivity series tend to react more vigorously, releasing more hydrogen gas and creating more fizz. For instance:

- **Zinc:** Reacts readily with dilute acids, producing noticeable fizz.
- **Magnesium:** Reacts even more vigorously, generating a significant amount of bubbles.
- **Aluminum:** Usually protected by a thin oxide layer, so reaction may be slower unless the oxide layer is removed or the acid is strong enough.

2. What safety precautions should be taken when exploring fizzy metals?

Handling reactive metals and acids requires proper safety measures, including:

- Wearing safety goggles to protect eyes from splashes or gas bubbles.
- Using gloves to prevent skin contact with acids and metals.
- Conducting experiments in a well-ventilated area to avoid accumulation of hydrogen gas, which is flammable.
- Using appropriate containers that can withstand chemical reactions.

Always follow safety guidelines and dispose of chemicals responsibly after experiments.

3. How can I demonstrate fizzy metal reactions at home or in the classroom?

Here's a simple experiment to observe fizzy reactions:

- Gather materials: zinc granules or strips, dilute hydrochloric acid (or vinegar as a safer alternative), a test tube or small container, safety goggles, and gloves.
- Place the zinc in the container.
- Carefully add the acid to the zinc.
- Observe the bubbling and fizzing as hydrogen gas is released.
- Optional: Capture the gas in a balloon attached to the opening to visualize hydrogen accumulation.

This experiment demonstrates the reaction vividly and helps learners understand gas evolution in

chemical reactions.

The Importance of Fizzy Metals in Science Education

Educational Value

Fizzy metals serve as an interactive teaching tool, making abstract chemical concepts tangible. They help students grasp:

- Reactivity series of metals
- Chemical reactions and gas evolution
- The properties of acids and bases
- Safety procedures in chemical experimentation

Real-World Applications

Understanding the reactions of fizzy metals has practical implications in various industries, including:

- Corrosion prevention and treatment
- Battery technology, where metal reactions generate electrical energy
- Hydrogen fuel development, utilizing hydrogen gas produced from metal reactions

Advanced Insights into Fizzy Metal Reactions

Factors Affecting Reaction Rate

The speed and intensity of fizzing depend on several factors:

- **Type of metal:** More reactive metals produce faster reactions.
- **Concentration of the acid:** Higher concentration increases reaction rate.
- **Surface area of the metal:** Finely divided metals react more quickly due to increased surface contact.
- **Temperature:** Elevated temperatures generally accelerate reactions.

Chemical Properties and Safety Concerns

Hydrogen gas, produced during these reactions, is highly flammable and explosive when mixed with air. Therefore, experiments should be conducted away from open flames or heat sources. Proper

ventilation is essential to prevent gas accumulation.

Conclusion

Exploring science fizzy metals and their reactions offers an exciting window into the world of chemistry. By understanding the underlying principles—such as metal reactivity, gas evolution, and acid-base interactions—students and enthusiasts can appreciate both the scientific concepts and practical applications. Whether conducting simple experiments at home or exploring advanced industrial processes, knowledge about fizzy metals enhances scientific literacy and fosters curiosity about the natural world.

Remember to always prioritize safety and proper handling when working with reactive metals and acids. With the right precautions, exploring fizzy metals can be a fun and educational journey into the fascinating realm of chemical reactions and gases.

Exploring Science Fizzy Metals 2 Answers: A Comprehensive Guide

When it comes to engaging science experiments that captivate both students and educators alike, science fizzy metals 2 answers stand out as a popular and educational activity. These experiments not only demonstrate fascinating chemical reactions but also serve as an excellent way to introduce concepts such as oxidation, acid-base reactions, and the properties of metals. Whether you're a teacher preparing a lesson plan, a student conducting a science project, or a science enthusiast exploring chemical reactions, understanding the core principles behind these fizzing metal experiments is essential. In this guide, we'll delve into the science behind fizzy metals, explore common "Science Fizzy Metals 2 Answers," and provide practical insights to deepen your understanding.

What Are Fizzing Metals in Science Experiments?

Fizzy metals experiments typically involve combining certain metals with acids or other reactive substances to produce gas bubbles—commonly carbon dioxide—that create a fizzing or bubbling effect. These reactions vividly illustrate how metals react with acids, release gases, and sometimes form new compounds. The term "fizzy" describes the effervescent nature of the reactions, which often look like bubbling or foaming.

Key Concepts:

- Chemical Reactivity of Metals: Not all metals react the same way; some are more reactive than others.
- Gas Production: The fizzing effect is due to the release of gases like hydrogen or carbon

dioxide.

- Reaction Types: Typically involve oxidation-reduction (redox) reactions, acid-metal reactions, or displacement reactions.

Common Metals Used in Fizzy Reactions

In science experiments, certain metals are favored because of their predictable reactions and safety profiles. Here's a list of frequently used metals:

- Zinc (Zn): Reacts with acids to produce hydrogen gas.
- Magnesium (Mg): Reacts vigorously with acids.
- Aluminum (Al): Reacts with acids but less vigorously.
- Iron (Fe): Reacts slowly with acids.
- Copper (Cu): Usually less reactive with acids.
- Sodium (Na), Potassium (K): Highly reactive, but less common in classroom experiments due to safety concerns.

The Core Science Behind Fizzy Metals Reactions

Understanding science fizzy metals 2 answers involves grasping the chemical reactions at play. The most common reaction in these experiments involves metals reacting with acids, such as hydrochloric acid (HCl) or vinegar (acetic acid).

General Reaction Formula:



For example, with zinc:



Here, hydrogen gas (H_2) is released, creating the fizzing or bubbling effect.

Why does this happen?

- The metal atoms lose electrons (oxidation) to form metal ions.
- Hydrogen ions from the acid gain electrons (reduction) to form hydrogen gas.
- The gas bubbles escape, causing the fizzing.

Exploring "Science Fizzy Metals 2 Answers"

The phrase "Science Fizzy Metals 2 Answers" suggests a reference to a specific set of questions or a quiz, often used in educational contexts. Typically, these are part of a science workbook or quiz aimed at testing understanding of chemical reactions involving metals. Common questions and their explanations include:

- Which metals produce the most fizzing when reacting with acids?

Answer:

Metals like magnesium and zinc tend to produce vigorous fizzing because of their higher reactivity. Alkali and alkaline earth metals (like sodium and calcium) react even more vigorously but are less common in classroom experiments due to safety concerns.

- What gas is produced during these reactions?

Answer:

Hydrogen gas (H_2) is the primary product when reactive metals react with acids. The bubbling or fizzing is due to hydrogen escaping from the solution.

Practical Experiment: Observing Fizzy Metal Reactions

Materials Needed:

- Metal strips (zinc, magnesium, aluminum)
- Hydrochloric acid or vinegar
- Test tubes or small containers
- Safety goggles and gloves
- Timer for measuring reaction duration

Procedure:

- Place a small piece of metal into the test tube.
- Carefully add acid to cover the metal.
- Observe the reaction? bubbles will form almost immediately.

- Record the intensity and duration of fizzing.
- Repeat with different metals to compare reactivity.

Safety Tip: Always wear proper safety gear and conduct experiments in a well-ventilated area.

Interpreting Results and Answering Common Questions

Q1: Why does zinc fizz more vigorously than aluminum?

A: Zinc is more reactive than aluminum, so it reacts faster and produces more hydrogen gas in a shorter time.

Q2: Can metals like copper produce fizzing in these reactions?

A: Copper is less reactive with acids, so its reaction with acids is slow or negligible, resulting in little to no fizzing.

Q3: How does surface area affect the reaction?

A: Increasing surface area (e.g., using powdered metal) speeds up the reaction, leading to more fizzing.

Broader Implications and Educational Value

Understanding these reactions helps students grasp fundamental chemical principles:

- Redox Reactions: Metals oxidize, and gases are reduced.
- Properties of Metals: Reactivity varies across the periodic table.
- Chemical Safety: Handling acids and reactive metals safely.

Educational Benefits:

- Visual demonstration of chemical reactions.
- Reinforcement of concepts like gas evolution and metal activity series.
- Development of scientific inquiry skills through observation and experimentation.

Additional Tips for Exploring Fizzy Metals

- Use different acids (e.g., vinegar, lemon juice) to compare reaction rates.
- Test multiple metals to understand reactivity trends.
- Keep detailed records of reaction times and fizz intensity.
- Incorporate real-world applications, such as metal corrosion or hydrogen fuel cells.

Conclusion

The science fizzy metals 2 answers encompass a fascinating area of chemistry that combines visual excitement with fundamental scientific principles. By exploring how different metals react with acids to produce gases like hydrogen, students and enthusiasts can deepen their understanding of chemical reactivity, the periodic table, and redox reactions. Whether conducting experiments, answering quiz questions, or simply observing the bubbling reactions, this knowledge opens doors to appreciating the dynamic and reactive nature of metals in our world. Always remember to prioritize safety and curiosity as you explore these fizzy, fascinating reactions!

Question What are fizzy metals in science experiments? Fizzy metals are metals that react with acids or other substances to produce bubbling or fizzing, often releasing hydrogen gas during the reaction. How can I safely explore fizzy metals at home? You can explore fizzy metals safely by using small pieces of reactive metals like zinc or magnesium with vinegar or lemon juice, conducting the experiment in a well-ventilated area and wearing safety goggles. Why do some metals fizz when reacting with acids? Metals fizz when reacting with acids because they release hydrogen gas as a byproduct, which forms bubbles and causes the fizzing effect. What are common examples of fizzy metals used in experiments? Common fizzy metals include zinc, magnesium, and aluminum, which readily react with acids to produce fizzing or bubbling. What science principles are demonstrated by exploring fizzy metals? Exploring fizzy metals demonstrates principles like chemical reactions, gas evolution, reactivity series of metals, and the conservation of mass during chemical reactions. Can fizzy metals be used to generate hydrogen fuel? Yes, reactive metals like zinc and magnesium can produce hydrogen gas when reacting with acids, which can be harnessed as a clean fuel source in controlled settings. What safety precautions should be taken when exploring fizzy metals? Always wear safety goggles, gloves, and work in a well-ventilated area. Avoid ingesting chemicals and handle acids carefully to prevent burns or reactions. How does the reactivity of metals influence their fizzing behavior? More reactive metals tend to fizz more vigorously because they react more readily with acids, releasing hydrogen gas quickly and producing more noticeable bubbling.

Related keywords: science experiments, fizzy metal reactions, chemical reactions, bubbling metals, science activities, educational experiments, metal oxidation, chemical reactions for kids, fizzy reactions, science exploration

FIZZY METALS 2 ANSWERS

fizzy metals 2 answers is a phrase that often arises in the context of chemistry riddles, puzzles, and educational quizzes designed to challenge students' understanding of chemical reactions and properties of metals. Whether you're a student preparing for exams, a teacher seeking engaging teaching tools, or simply a curious learner exploring the fascinating world of chemistry, understanding the concept of "fizzy metals" can deepen your appreciation for how metals interact with acids and other compounds to produce effervescent reactions. In this comprehensive guide, we will explore the nature of fizzy metals, provide the most common answers to related questions, and delve into the science behind these bubbly reactions.

Understanding Fizzy Metals: What Are They?

Definition and Characteristics

Fizzy metals refer to metals that, when reacted with acids, produce a bubbling or fizzing effect due to the release of gases such as hydrogen. This effervescent reaction is a hallmark of metals that are reactive enough to displace hydrogen from acids like hydrochloric acid (HCl) or sulfuric acid (H₂SO₄). The fizzing is essentially the rapid formation of hydrogen gas bubbles escaping from the reaction mixture.

Key characteristics of fizzy metals include:

- Reactivity with acids
- Production of hydrogen gas
- Often used in educational demonstrations
- Found on the reactivity series of metals, especially those higher up (e.g., zinc, magnesium, aluminum)

Common Metals That Are Considered "Fizzy"

Among metals commonly associated with fizzy reactions, the following are notable:

- Zinc (Zn)
- Magnesium (Mg)
- Aluminum (Al)
- Iron (Fe)
- Calcium (Ca)

- Lithium (Li)

While some of these metals react vigorously, others do so more mildly. Their position on the reactivity series determines their ability to produce fizzing reactions with acids.

Top Two Answers to "Fizzy Metals" Questions

Many questions related to "fizzy metals" tend to focus on identifying which metals produce fizzing reactions and understanding the science behind these reactions. Let's examine the two most common answers that frequently appear in educational contexts.

1. Zinc (Zn)

Zinc is perhaps the most frequently cited "fizzy metal" in chemistry exercises. It reacts readily with dilute acids, especially hydrochloric acid, to produce hydrogen gas, which causes bubbling or fizzing.

Reaction Example:



Why Zinc?

- Moderate reactivity: Reacts readily but safely in lab conditions.
- Commonly used in classroom experiments to demonstrate gas evolution.
- Produces noticeable fizzing, making it ideal for educational demonstrations.

Applications:

- Galvanization of steel
- Batteries (zinc-carbon batteries)
- Educational demonstrations for learning about gas evolution and metal reactivity.

2. Magnesium (Mg)

Another prominent "fizzy metal" answer is magnesium, known for its vigorous reaction with acids and its bright, shiny appearance.

Reaction Example:



Why Magnesium?

- Highly reactive: Produces a rapid and vigorous fizzing.
- Used in fireworks and flares due to bright combustion.
- Often included in experiments to illustrate reactive metals.

Educational Significance:

- Demonstrates faster hydrogen evolution than zinc.
- Shows the reactivity trend in the alkali and alkaline earth metals.

Exploring Other Metals and Their Fizzing Reactions

While zinc and magnesium are the most common answers, other metals can also produce fizzing reactions under certain conditions.

Aluminum

- Reacts with acids but often requires an activator (like mercury or a strong acid concentration) to overcome its oxide layer.
- When reacted, produces hydrogen gas with bubbling.

Iron

- Reacts slowly with acids, producing less vigorous fizzing.
- Used in corrosion studies and rust experiments.

Calcium and Lithium

- More reactive alkaline earth and alkali metals.
- Lithium reacts rapidly with acids, producing noticeable fizzing.
- Calcium produces fizzing but less vigorously compared to magnesium or zinc.

Factors Influencing Fizzing Reactions

Several factors determine how vigorously a metal reacts with acids and produces fizzing:

- **Reactivity Series Position:** Metals higher in the reactivity series (like lithium, potassium) react more vigorously.
- **Surface Area:** Finely divided metals react faster due to increased surface exposure.
- **Type of Acid:** Strong acids like hydrochloric acid produce more vigorous reactions than weaker acids.
- **Temperature:** Higher temperatures speed up the reaction rate, increasing fizzing.
- **Presence of Oxide Layers:** Metals like aluminum are coated with oxide layers that slow down reactions unless removed or bypassed.

Practical Applications and Demonstrations

Understanding fizzy metals isn't just theoretical; it has practical applications in industry, education, and scientific research.

Educational Demonstrations

- Showing gas evolution with zinc or magnesium in class.
- Comparing the reactivity of different metals.
- Visualizing the reactivity series in a tangible way.

Industrial Uses

- Hydrogen production: Metals reacting with acids to generate hydrogen gas for fuel or chemical synthesis.
- Metal corrosion studies: Observing how metals react with acids and other corrosive agents.
- Battery technology: Utilizing zinc and magnesium in primary and secondary batteries.

Conclusion: The Science Behind Fizzy Metals

The phrase "fizzy metals" encapsulates a fascinating aspect of chemistry—the reaction of certain metals with acids to produce hydrogen gas, resulting in bubbling or fizzing. Zinc and magnesium are the top two answers often associated with this phenomenon, thanks to their well-understood reactions and safety in educational settings. Recognizing these reactions not only helps in answering quiz questions but also enhances understanding of chemical reactivity, the behavior of metals, and their applications across various fields. Whether you're conducting simple classroom experiments or exploring advanced industrial processes, the principles behind fizzy metals serve as

a fundamental building block in the study of chemistry.

Remember: Always handle reactive metals and acids with appropriate safety precautions, including wearing gloves, goggles, and working in a well-ventilated area.

Fizzy Metals 2 Answers: A Comprehensive Review of the Popular Puzzle Game

Puzzle games have long captivated players with their blend of logic, creativity, and strategic thinking. Among these, Fizzy Metals 2 Answers has emerged as a standout title, captivating a broad audience with its engaging gameplay and challenging puzzles. This article provides an in-depth analysis of the game, exploring its core features, gameplay mechanics, solutions, and overall appeal, to help both new players and seasoned enthusiasts understand what makes Fizzy Metals 2 a noteworthy addition to the puzzle genre.

Introduction to Fizzy Metals 2

Fizzy Metals 2 Answers is a sequel to the original Fizzy Metals game, developed by a team dedicated to creating brain-teasing puzzles that test players' problem-solving skills. The game involves manipulating a series of metallic objects—such as balls, wires, and other components—to achieve specific objectives, often involving guiding metals through complex pathways, avoiding obstacles, or activating switches.

The core premise combines physics-based mechanics with strategic planning, requiring players to think ahead and experiment with various configurations. Its popularity stems from its intuitive interface, progressively challenging levels, and the satisfaction derived from solving intricate puzzles.

Gameplay Mechanics and Features

Core Gameplay Elements

At the heart of Fizzy Metals 2 are puzzles that revolve around guiding metallic elements from a starting point to a designated target. Players interact with a variety of objects, including:

- Balls and metallic spheres
- Wires, tubes, and pathways
- Switches and levers
- Obstacles like barriers or moving parts

The game leverages realistic physics, making the behavior of metals, gravity, and other forces

believable and engaging.

Level Design and Progression

The game features numerous levels, each increasing in complexity. Early levels serve as tutorials, introducing basic mechanics, while later stages challenge players with multi-step solutions and more intricate setups.

Levels are designed with multiple solutions in mind, encouraging experimentation. Some puzzles can be solved with simple moves, while others require creative combinations and precise timing.

User Interface and Controls

Fizzy Metals 2 boasts a clean, user-friendly interface that allows players to easily manipulate objects with simple touch or mouse controls. Features include:

- Drag-and-drop mechanics for placing and moving objects
- Undo and reset options for experimentation
- Hint system to assist players stuck on difficult puzzles

Additional Features

- Hint system: Offers clues to help players progress without spoiling the solution.
- Level editor: Some versions include tools for creating custom puzzles, fostering a community of puzzle enthusiasts.
- Achievements and rewards: Incentivize players to complete levels efficiently or discover alternative solutions.

Solutions and Answers to Common Puzzles

One of the most sought-after aspects of Fizzy Metals 2 is its challenging puzzles, prompting players to seek solutions or hints. While the game encourages experimentation, many players prefer to see step-by-step answers for difficult levels.

General Approach to Solving Puzzles

Before delving into specific solutions, it's helpful to understand the general strategies:

- Analyze the level layout: Identify start points, endpoints, obstacles, and interactable objects.
- Plan your moves: Think ahead about the sequence of actions needed.
- Experiment: Use trial and error to see how objects behave under different configurations.
- Utilize hints: When stuck, hints can guide you toward the correct approach.

Example Solutions for Selected Levels

Below are summarized solutions for a few representative levels. For detailed walkthroughs, players often consult dedicated forums or walkthrough videos.

Level 5: "Metal Bridge"

Objective: Guide the metal ball across a set of obstacles to the target.

Solution overview:

- Place a wire to create a pathway that redirects the ball.
- Use a switch to lift barriers temporarily.
- Drop the ball from a high point to gain momentum.
- Carefully time the activation of switches to allow the ball to pass through.

Level 12: "The Collapsing Tower"

Objective: Prevent the tower of metallic blocks from collapsing while guiding the ball through.

Solution overview:

- Remove certain supporting blocks to control the collapse direction.
- Use wires to redirect falling metals safely.
- Activate switches to stabilize critical parts temporarily.
- Guide the metal ball to the target after ensuring the tower's controlled collapse.

Tips for Finding Your Own Solutions

- Break down complex puzzles into smaller parts.
- Experiment with different object placements.
- Watch gameplay videos for inspiration.
- Keep notes on what strategies work best.

Pros and Cons of Fizzy Metals 2

Pros

- Engaging physics-based puzzles that challenge players' strategic thinking.
- Progressively difficult levels suited for all skill levels.
- Intuitive controls make gameplay accessible.
- Multiple solutions encourage creativity.
- Community resources and hints assist players in difficult puzzles.
- Beautiful, minimalist design enhances user experience.

Cons

- Some puzzles can become repetitive after extended play.
- Limited tutorial content may pose initial challenges for new players.
- The level editor feature is not always available or fully functional in all versions.
- Certain solutions may require trial and error, which can be frustrating.
- Hint system might reveal too much, reducing the puzzle-solving satisfaction for some users.

Final Thoughts and Recommendations

Fizzy Metals 2 Answers offers a compelling blend of physics-based puzzle mechanics and strategic gameplay that appeals to puzzle enthusiasts and casual players alike. Its well-designed levels, combined with the freedom to experiment and discover multiple solutions, make it a rewarding experience for those who enjoy mental challenges.

For players seeking to improve their problem-solving skills or simply enjoy engaging puzzles, Fizzy Metals 2 is a highly recommended title. While it may present some initial learning curves and occasional frustration, the satisfaction of solving complex puzzles and mastering the mechanics outweighs these minor drawbacks.

Tips for New Players

- Start with the early levels to familiarize yourself with mechanics.
- Use hints sparingly to maintain the challenge.
- Experiment with different configurations for tricky puzzles.
- Explore community forums for solutions and creative ideas.
- Take breaks to avoid frustration and approach puzzles with a fresh perspective.

In conclusion, Fizzy Metals 2 Answers stands out as a thoughtfully designed puzzle game that successfully combines realistic physics with engaging gameplay. Its depth, variety, and challenge make it a valuable addition to any puzzle game collection. Whether you're a casual player or a dedicated puzzle solver, the game offers hours of stimulating entertainment that will keep your mind sharp and your curiosity piqued.

Question What are fizzy metals in chemistry, and how do they react? Fizzy metals typically refer to reactive metals like sodium or magnesium that release gas bubbles when they react with acids or water, producing fizzing or bubbling due to hydrogen gas formation. **Why do some metals produce a fizzy or bubbling reaction in experiments?** Metals like magnesium and zinc produce fizzing when they react with acids because they release hydrogen gas, which forms bubbles and causes the fizzy appearance during the chemical reaction. **How can fizzy metals be used in practical applications?** Fizzy metals are used in medical imaging (e.g., magnesium in biodegradable implants), fireworks, and chemical reactions where controlled gas release is needed, leveraging their bubbling reactions. **Are all reactive metals considered fizzy metals?** Not all reactive metals are called fizzy metals; the term often informally refers to metals that produce noticeable fizzing or bubbling reactions, such as magnesium, zinc, or aluminum when reacting with acids. **What safety precautions should be taken when handling fizzy metals?** Handling fizzy metals requires gloves and eye protection, as reactions can be vigorous and produce hydrogen gas, which is flammable. Work in a well-ventilated area and follow proper chemical safety protocols. **Can the fizzing reaction of metals be used to identify them?** Yes, observing the fizzing or bubbling reaction when metals react with acids can help identify reactive metals like magnesium or zinc, as different metals produce characteristic reaction rates and gas volumes.

Related keywords: carbonated metals, metallic effervescence, bubbly metals, fizzy metal reactions, metallic foaming, effervescent alloys, bubbling metals, reactive metallic solutions, metallic carbonation, fizzy metal compounds

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